

therefore suggest that the lengthening of the C(1)–C(7) bond is caused by this rotation, which has decreased the overlap between the two p_z orbitals and therefore decreased the double-bond character of the bond.

The molecule is distorted from planarity by an unusual amount. Omitting the 'outside' atoms of the isobutyl group, C(10), C(11), C(12), the distortion of one half of the molecule from complete planarity may be approximately described by a rotation of the (nearly) planar group C(1), C(7), C(8), N, C(9), Cu through an angle of 25° about the line C(1)–C(9). Equations for various least-squares planes and the distances of the several atoms from these planes are given in Table 4. Comparison of these distances with those found in other compounds shows a number of differences. The metal atom is farther (1.12 \AA) out of the plane defined by the benzene ring, the farthest heretofore reported being 0.75 \AA in bis-(*N*-*t*-butylsalicylaldiminato)palladium(II) (Day, Glick & Hoard, 1968). The nitrogen atom is also farther (0.25 \AA) out of the plane and is out in the same direction as the metal, whereas in other compounds it is out in the opposite direction from the metal atom (*cf.* Table 6, Jain & Lingafelter, 1967). The dihedral angle between the coordination plane CuN_2O_2 and the mean plane of the *o*-hydroxyacetophenone imine residue is 136.6° , which may be compared with the corresponding angle of 144.4° in bis-(*N*-*t*-butylsalicylaldiminato)palladium.

The packing of the molecules may be seen in Fig. 2. None of the intermolecular contact distances is unusual.

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References

- ALEXANDER, L. E. & SMITH, G. S. (1964). *Acta Cryst.* **17**, 1195.
 CROMER, D. & WABER, J. (1965). *Acta Cryst.* **18**, 104.
 DAY, V. W., GLICK, M. D. & HOARD, J. L. (1968). *J. Amer. Chem. Soc.* **90**, 4803.
 HALL, D., SHEAT, S. V. & WATERS, T. N. (1968). *J. Chem. Soc. A* **1968**, 460.
 JAIN, P. C. & LINGAFELTER, E. C. (1967). *Acta Cryst.* **23**, 127.
 LINGAFELTER, E. C. & BRAUN, R. L. (1966). *J. Amer. Chem. Soc.* **88**, 2951.
 LINGAFELTER, E. C., SIMMONS, G. L., MOROSIN, B., SCHERINGER, C. & FREIBURG, C. (1961). *Acta Cryst.* **14**, 1222.
 STEWART, J. M. (1964). *Crystal Structure Calculation System for the IBM 709, 7090, 7094*. Technical Report TR-64-6, Computer Science Center, Univ. of Maryland and Research Computer Laboratory, Univ. of Washington.
 STEWART, R. F., DAVIDSON, E. R. & SIMPSON, W. T. (1965). *J. Chem. Phys.* **42**, 3175.
 WATERS, T. N. & HALL, D. (1959). *J. Chem. Soc.* p. 1200.
 WEI, L., STOGSDILL, R. M. & LINGAFELTER, E. C. (1964). *Acta Cryst.* **17**, 1058.

Acta Cryst. (1970). B26, 1812

The Crystal Structure of Compounds with (N-P)_n Rings. VII.* Refinement of the Crystal Structure of Hexabromocyclotriphosphazene, N₃P₃Br₆

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The crystal structure of N₃P₃Br₆, as determined by Giglio & Puliti (1967), has been refined from three-dimensional intensity data, collected with Mo radiation on an automatic three-circle diffractometer. Two crystals were used. Data processing and refinement were carried out independently with the data obtained from each crystal. The intensities were corrected for absorption and during the refinement a correction for extinction was applied according to Zachariasen (1967, 1968). Atomic coordinates obtained from the two crystals agree within the standard deviations. The molecules N₃P₃Br₆ lie on crystallographic mirror planes of the space group *Pnma*, with unit-cell dimensions $a=14.463$ (0.002), $b=13.410$ (0.003) and $c=6.601$ (0.001) Å. There are four short intermolecular Br...Br distances (3.65–3.85 Å) and a short N...Br distance (3.24 Å). The two independent angles P–N–P of a molecule are significantly different: 119.3 (0.6) and 122.4 (0.5)°. Other chemically equivalent bond lengths and angles are equal within experimental error. Mean values (and individual standard deviations) are P–N 1.576 (0.008), P–Br 2.162 (0.004) Å, N–P–N 118.5 (0.5) and Br–P–Br 102.1 (0.1)°. The (N–P)₃ ring is non-planar and slightly chair-shaped. The dihedral angles of the ring bonds are -6.7 , 6.3 and -5.5° and the largest distance from a ring atom to the least-squares plane through the ring is 0.047 \AA .

Introduction

The crystal structures of several chloro- and fluorocyclophosphazenes are known with some precision,

viz. N₃P₃Cl₆ (Wilson & Carroll, 1960), N₄P₄Cl₈ [two crystal modifications, *K* form (Hazekamp, Migchelsen & Vos, 1962) and *T* form (Wagner & Vos, 1968)], N₅P₅Cl₁₀ (Schlueter & Jacobson, 1968), N₃P₃F₆ (Dougill, 1963) and N₄P₄F₈ (McGeachin & Tromans, 1961).

* Part VI: Olthof (1969).

As to the bromo compounds the only crystal structure known is that of the trimer $N_3P_3Br_6$. This structure was determined by Giglio & Puliti (1967) from the intensities of 538 X-ray reflexions, estimated visually on Weissenberg films. The accuracy of the results is not high, e.g. the values found for the P-Br bond lengths range from 2.14 to 2.20 Å with an individual standard deviation of 0.015 Å.

We intend to investigate the crystal structures of the homologous bromocyclophosphazenes in order to elucidate the conformations of the phosphazene rings. Refinement of the crystal structure of the trimer $N_3P_3Br_6$, from new experimental data, seemed necessary, if reliable comparisons with the structures of related compounds are to be made.

Since absorption effects of X-rays in crystals of the bromocyclophosphazenes will be very substantial (linear absorption coefficient for Mo radiation more than 200 cm^{-1}), we thought it worth while to use two independent sets of intensities, obtained from two different crystals, for the refinement of the crystal structure of $N_3P_3Br_6$, in order to see whether the determined atomic coordinates would be essentially the same.

The results of the refinement of $N_3P_3Br_6$ are given in the present paper. The crystal structure of the tetramer $N_4P_4Br_8$, which has been determined by us, will be described in a following paper.

Experimental

The substance was prepared from phosphorus(III) bromide, ammonium bromide and bromine according to the procedure of Coxon, Sowerby & Tranter (1965). From the reaction mixture, containing trimeric and tetrameric bromocyclophosphazene and possibly higher homologues, suitable crystals of $N_3P_3Br_6$ were obtained by fractional crystallization from petroleum.

The crystals are orthorhombic. Weissenberg photographs showed the systematic absences $0kl$ for $k+l=2n+1$ and $hk0$ for $h=2n+1$. The space group therefore is $Pnma$ (no. 62) or $Pn2_1a$ (no. 33). The unit cell contains four molecules $N_3P_3Br_6$. In the refinement it was assumed that $Pnma$ is the correct space group, implying that the molecules have a crystallographic mirror plane.

Accurate unit-cell dimensions were determined from zero-level Weissenberg photographs around the b and c axes. Reflexions from a NaCl crystal were superposed on the films for calibration. The lengths of the unit-cell edges are:

Present work		Giglio & Puliti (1967)	
$a=14.463$, s.d. 0.002 Å		14.43, s.d. 0.02 Å	
$b=13.410$	0.003	13.36	0.02
$c=6.601$	0.001	6.63	0.01

Intensities were measured at room temperature on a single-crystal three-circle automatic Nonius diffractometer by the $\theta-2\theta$ scan method. All 1176 independent reflexions up to $\sin \theta/\lambda=0.60\text{ \AA}^{-1}$ were measured using Zr filtered Mo radiation. The intensities of all reflexions were measured twice; two different crystals (I and II) of approximate dimensions $0.1 \times 0.1 \times 0.2$ mm and $0.2 \times 0.2 \times 0.3$ mm respectively were used. Crystal I gave 995 non-zero reflexions and crystal II 1004. The data for each crystal were processed separately and the refinement was carried out with each series of F_o values independently.

The intensities were corrected for Lorentz and polarization effects and for absorption. The latter corrections were calculated with a computer program based on the Busing & Levy (1957) scheme. The linear absorption coefficient for Mo radiation is very high, viz. 203 cm^{-1} . The transmission factor $(1/V) \int \exp(-\mu d)dV$ ranged from 0.10 to 0.14 for I and from 0.02 to 0.06 for II. After the absorption correction the F_o values of the two crystals were compared. After scaling the factor $R' = \sum ||F_{II}| - |F_I|| / \sum |F_I|$ was 10.6% for the 899 common reflexions. For the stronger reflexions the scaled F_o values of II were generally smaller than those of I. As will be seen later this behaviour can only partly be ascribed to extinction effects.

Refinement

Least-squares refinements were carried out based on the F_o values of each crystal separately. A least-squares program working according to the block-diagonal approximation (Cruickshank, 1961) was used. The atomic coordinates and thermal parameters from the paper by Giglio & Puliti (1967) were taken as a

Table 1. *Final fractional coordinates*

Standard deviations are in parentheses.

	Crystal I			Crystal II		
	x	y	z	x	y	z
N(1)	0.1430 (10)	0.25	0.5967 (30)	0.1434 (11)	0.25	0.5940 (29)
N(2)	-0.0077 (7)	0.1499 (8)	0.4566 (19)	-0.0081 (8)	0.1507 (9)	0.4528 (24)
P(1)	-0.0605 (3)	0.25	0.4121 (9)	-0.0607 (4)	0.25	0.4117 (9)
P(2)	0.0907 (2)	0.1477 (3)	0.5556 (6)	0.0910 (3)	0.1480 (3)	0.5554 (6)
Br(1)	-0.1895 (2)	0.25	0.5783 (4)	-0.1895 (2)	0.25	0.5789 (4)
Br(2)	-0.1078 (2)	0.25	0.1015 (4)	-0.1078 (2)	0.25	0.1019 (4)
Br(3)	0.0823 (1)	0.0714 (1)	0.8445 (3)	0.0824 (1)	0.0713 (1)	0.8442 (3)
Br(4)	0.1806 (1)	0.0507 (1)	0.3845 (3)	0.1806 (1)	0.0505 (1)	0.3844 (3)

starting point. Atomic scattering factors were taken from Doyle & Turner (1968).

In the initial cycles the scale factor, the atomic coordinates and anisotropic thermal parameters were refined without accounting for anomalous scattering effects and with uniform weights. Later the anomalous scattering by phosphorus and bromine (*International Tables for X-ray Crystallography*, 1962) was incorporated and a weighting scheme of the form $w = \{w_c^{-1} + p|F_o|^2\}^{-1}$ was introduced. The term w_c is the weight from counting statistics and p is a constant, chosen so as to make $(\Delta F)^2$, averaged over groups of reflexions, as far as possible independent on $|F_o|$. The value 4×10^{-4} was given to p for crystal I and 9×10^{-4} for II. Furthermore, the extinction parameter p_{ex} was refined according to Zachariasen's (1967, 1968) method. This parameter occurs in the function M , which is to be minimized, as follows:

$$M = \sum w [|F_o| - K|F_c| \{1 + C(\theta)p_{ex}|F_c|^2\}^{-1/4}]^2$$

$$\text{with } C(\theta) = 10^{-2} (p_2/p_1) (1/\sin 2\theta) \bar{T}$$

$$\text{and } p_{ex} = 2 \times 10^2 (e^2/mc^2)^2 (\lambda^2/V^2) r^*$$

(See Zachariasen (1968) for the meaning of the symbols.)

At the end of the refinement the indices $R = \{\sum |\Delta F|^2 / \sum |F_o|^2\}^{1/2}$ and $R_w = \{\sum w |\Delta F|^2 / \sum w |F_o|^2\}^{1/2}$ were 9.6 and 6.5% for crystal I and 9.4 and 8.0% for II. The theoretical values for R and R_w based on counting statistics are lower, *viz.* 8.5 and 4.2% for I, and 7.6 and 5.2% for II. The differences between F_o and F_c are thus larger than could be expected on the basis of counting statistics alone, indicating that there are additional errors in the F_o values. Presumably these errors are mainly due to the difficulties encountered in calculating the absorption corrections. The refined

value of p_{ex} is 0.005 for I and 0.014 for II. With these numbers the extinction corrections $y^{1/2} = \{1 + C(\theta)p_{ex}|F_c|^2\}^{-1/4}$ are calculated as about 0.95 and 0.85 for the strong reflexions of I and II respectively. It thus appeared that the extinction effects are rather small in both crystals, though somewhat larger in the bigger one.

The last four cycles of the least-squares refinement were also run without applying extinction corrections. It appeared that this brought about only slight changes in the atomic coordinates, in no case larger than the standard deviations. Furthermore, neglect of the extinction corrections caused only moderate changes in the thermal parameters, the scale factor, the standard deviations and the indices R and R_w . The sense of these changes is as expected: compared with the refinement with extinction corrections the thermal parameters and scale factor decreased, the standard deviations in all parameters increased and R increased (1% for I and 4% for II). The change in R_w was less than 1%.

Results and discussion

The final coordinates with their standard deviations calculated by the least-squares program are given in Table 1. It is seen that corresponding coordinates, determined from crystal I and crystal II, are in most cases equal within the standard deviations. The anisotropic temperature factor parameters and their standard deviations are given in Table 2. Here again corresponding values obtained for the two crystals agree within experimental error, indicating that the two sets of F_o values (after correction for extinction) do not show significant differences in their variation with $\sin \theta/\lambda$. Observed and calculated structure factors are compared in Table 3. In this Table the extinction correction is applied to the observed amplitudes F_o .

Table 2. Parameters ($\text{\AA}^2 \times 10^{-4}$) of the anisotropic temperature factors $\exp[-2\pi^2(h^2a^{*2}U_{11} + \dots + 2hka^*b^*U_{12} + \dots)]$

Standard deviations are in parentheses.

	Crystal I					
	U_{11}	U_{22}	U_{33}	$2U_{12}$	$2U_{23}$	$2U_{13}$
N(1)	257 (94)	345 (99)	929 (160)	0	0	-526 (217)
N(2)	349 (65)	262 (64)	652 (90)	-84 (110)	107 (132)	-262 (131)
P(1)	222 (27)	386 (31)	439 (34)	0	0	-98 (54)
P(2)	268 (19)	290 (20)	496 (25)	64 (34)	1 (38)	-61 (36)
Br(1)	375 (12)	615 (15)	620 (16)	0	0	233 (26)
Br(2)	555 (15)	752 (17)	441 (14)	0	0	-224 (26)
Br(3)	669 (12)	655 (11)	519 (10)	-6 (19)	274 (19)	9 (19)
Br(4)	509 (9)	514 (9)	702 (12)	197 (17)	-216 (20)	271 (19)
	Crystal II					
N(1)	205 (88)	409 (101)	610 (130)	0	0	-439 (184)
N(2)	326 (70)	337 (72)	952 (121)	-88 (123)	-168 (165)	-234 (156)
P(1)	201 (26)	389 (30)	458 (34)	0	0	-131 (52)
P(2)	274 (19)	278 (19)	462 (23)	48 (35)	18 (38)	-79 (35)
Br(1)	381 (13)	592 (15)	608 (16)	0	0	208 (25)
Br(2)	587 (16)	755 (18)	408 (14)	0	0	-155 (25)
Br(3)	692 (12)	629 (11)	526 (10)	9 (20)	289 (20)	38 (19)
Br(4)	506 (10)	513 (10)	728 (12)	230 (17)	-196 (21)	238 (19)

A projection of the crystal structure along the c axis is given in Fig. 1. The (N-P)3 rings of the molecules A and C lie approximately at height z=0.5, and those

of the molecules B and D approximately at height z=0. The numbering of the atoms in molecule A corresponds to that used in the Tables. Fig. 2 shows a

Table 3. Observed and calculated structure factors

Table with 48 columns (H, K, L, FO, FC, AC, BC, etc.) and 100 rows of numerical data representing observed and calculated structure factors for various crystallographic indices.

Table 3 (cont.)

Crystal I

Crystal I																																										
H	K	L	FO	FC	AC	BC	H	K	L	FO	FC	H	K	L	FO	FC	AC	BC	H	K	L	FO	FC	H	K	L	FO	FC	AC	BC	H	K	L	FO	FC	AC	BC					
9	5	5	109	180	177	32	0	9	5	628	605	601	74	9	0	6	459	400	398	44	7	3	6	134	104	103	11	0	8	6	793	744	-738	-93	7	1	7	367	330	327	42	
10	5	5	280	139	-139	-9	1	9	5	184	119	-118	-17	10	0	6	234	254	-293	+42	9	6	6	106	30	37	10	1	8	6	182	203	200	33	1	2	7	137	39	35	-7	
11	5	5	240	172	-172	-10	2	9	5	172	112	-111	17	1	1	6	85	131	-131	-7	10	3	0	178	48	45	17	2	8	6	239	230	-228	-24	2	2	7	473	511	505	73	
12	5	5	286	259	-249	-26	3	9	5	82	196	-194	-33	2	1	5	659	629	-623	-86	0	4	0	684	696	-690	-91	3	8	6	167	146	-145	-16	3	2	7	238	190	187	31	
2	6	5	654	695	-689	74	4	9	5	325	341	-339	-29	4	1	6	249	448	-457	-17	4	6	6	180	184	151	30	4	8	6	188	309	305	44	4	2	7	162	98	97	11	
3	6	5	52	127	125	22	6	9	5	397	386	-384	-36	5	1	5	82	60	-61	-30	5	4	0	297	279	277	32	5	8	6	203	43	-42	-5	5	2	7	357	325	-320	-56	
4	6	5	602	542	-539	90	8	9	5	103	184	-183	-24	8	1	6	135	88	-88	0	7	4	0	80	43	-42	4	6	8	6	391	336	333	44	7	2	7	96	70	-67	-19	
5	6	5	281	560	-541	-77	9	9	5	147	43	-43	-4	7	1	5	328	339	-337	31	8	4	0	302	269	267	31	1	6	6	219	104	101	25	0	3	7	221	197	194	32	
6	6	5	225	20	-19	6	1	10	5	283	273	-272	-26	8	1	6	334	222	-218	39	9	4	0	146	189	-188	-19	3	9	6	104	50	50	3	1	3	7	257	247	-245	-31	
7	6	5	52	212	211	26	2	10	5	186	224	-224	-7	10	1	5	184	188	-186	26	10	4	6	30	135	-149	-42	4	9	6	154	126	124	23	2	3	7	200	148	-147	-17	
8	6	5	502	540	-537	90	3	10	5	275	325	-322	49	11	1	5	195	193	-189	-21	1	5	6	188	40	-37	-12	5	9	6	189	109	-103	-35	3	3	7	366	304	-302	-33	
9	6	5	418	400	-397	47	4	10	5	258	107	-107	7	0	2	6	1003	960	-954	108	2	5	0	747	782	-776	-91	0	10	6	391	377	375	39	4	3	7	120	33	31	11	
10	6	5	181	152	-180	-26	5	10	5	309	352	-347	-62	1	2	5	209	261	-259	-37	3	5	6	233	199	199	40	1	10	6	140	293	270	39	5	3	7	216	246	-244	-28	
1	7	5	520	457	-453	-58	6	10	5	151	57	57	7	2	2	6	711	288	-287	28	4	4	0	544	559	-555	-56	2	10	6	201	87	-83	-25	7	3	7	141	62	62	6	
2	7	5	240	120	119	12	7	10	5	238	254	-252	29	3	2	5	64	209	-208	25	5	5	0	551	589	-585	-59	1	0	7	473	467	465	46	1	4	7	284	122	-121	-14	
4	7	5	68	144	-143	-14	0	11	5	100	28	26	11	4	2	6	401	437	-443	-54	8	5	6	361	393	-390	50	2	0	7	384	247	245	30	3	4	7	185	103	-101	-18	
6	7	5	184	79	-79	1	1	11	5	139	84	83	13	6	2	5	429	463	-460	-55	9	5	0	363	330	-327	46	4	0	7	307	182	-180	-31	4	4	7	403	307	-305	-33	
7	7	5	698	609	-608	-94	2	11	5	327	315	-312	46	7	2	5	429	453	-449	-60	1	6	0	259	219	219	41	5	0	7	263	230	-230	-20	5	4	7	336	325	322	44	
9	7	5	407	389	-382	-51	3	11	5	197	159	-156	29	8	2	5	365	213	-212	-19	2	6	0	304	242	-239	-39	6	0	7	290	261	256	47	0	5	7	583	557	550	67	
1	8	5	178	241	-236	-49	4	11	5	254	15	-14	9	1	6	9	122	169	-160	42	7	0	7	122	169	-160	42	7	0	7	490	424	-421	-52	1	5	7	97	154	154	14	
2	8	5	177	227	227	17	5	11	5	108	24	-23	6	10	2	6	270	193	-190	39	8	6	0	295	273	270	39	0	1	7	535	467	463	59	3	5	7	205	136	-130	-42	
3	8	5	234	264	-262	34	0	0	6	364	420	-419	-25	1	3	6	959	957	-952	-72	8	6	0	105	29	-26	9	1	7	7	60	49	41	10	4	5	7	150	152	-147	-37	
4	8	5	856	839	-833	-105	1	0	0	782	756	-751	-69	2	3	6	245	155	-154	9	9	6	0	251	69	-67	-24	2	1	7	198	124	123	12	1	6	7	148	190	-187	-32	
5	8	5	250	157	-155	24	2	0	6	367	399	-393	69	3	3	5	452	450	-446	56	2	7	6	388	423	-420	53	3	1	7	346	333	-330	-45	2	6	7	238	105	105	3	
8	8	5	400	387	-384	-50	3	0	6	474	576	-577	-57	4	3	5	419	477	-469	-84	3	6	0	249	318	-312	41	4	1	7	323	298	-296	-36	0	7	7	77	65	65	18	
9	8	5	188	192	-191	-19	0	0	6	323	360	-354	-69	5	3	5	322	239	-238	23	4	7	6	134	170	-172	17	5	1	7	75	133	-131	-24	1	7	7	336	46	-45	-7	
10	8	5	169	119	-119	-19	8	0	6	157	167	-167	11	6	3	5	150	179	-178	15	8	7	0	99	110	-109	-15	6	1	7	315	307	-305	-31								

projection along the b axis of the molecules lying on the mirror plane $y=0.25$. All independent intermolecular distances which are nearly equal to or shorter than the sum of the corresponding van der Waals radii are given in the Figures (Pauling, 1960; $r(\text{N})=1.5$, $r(\text{P})=1.9$, $r(\text{Br})=1.95 \text{ \AA}$).

It is interesting to note that in the crystal structure of $\text{N}_3\text{P}_3\text{Cl}_6$ (Wilson & Carroll, 1960) only the intermolecular distances of the type given in Fig. 2 are shorter than the corresponding van der Waals distances, where-

as in the crystal structure of $\text{N}_3\text{P}_3\text{F}_6$ (Dougill, 1963) no short intermolecular distances occur.

The short distance $\text{N}(1)\cdots\text{Br}(1)$, shown in Fig. 2, is part of the approximately linear array $\text{N}(1)\cdots\text{Br}(1)\cdots\text{P}(1)$. Likewise the short distance $\text{Br}(1)\cdots\text{Br}(2)$ is part of the linear array $\text{Br}(1)\cdots\text{Br}(2)\cdots\text{P}(1, B)$. Such linear arrays have been observed in the crystal structures of polyhalogen compounds and have there been interpreted in terms of three-centre four-electron bonding (Wiebenga &

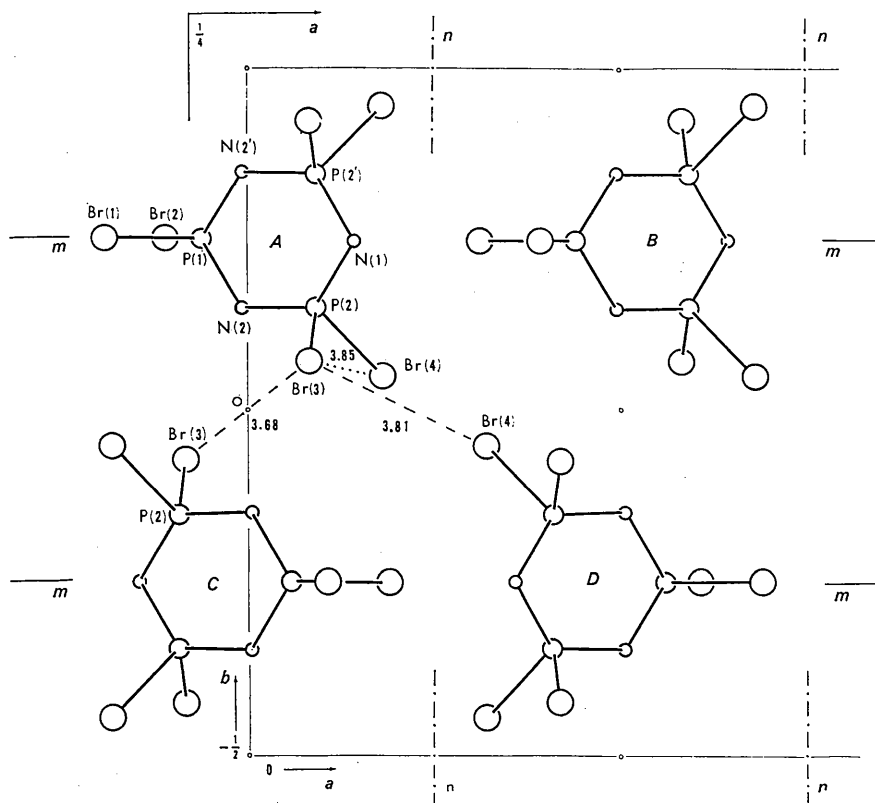


Fig. 1. Projection of the crystal structure along the c axis. The molecules lie on mirror planes at $y=0.25$ and $y=0.75$. The rings of the molecules A and C are approximately at height $z=0.5$, and those of B and D approximately at height $z=0$. The short intermolecular distances are as follows: $\text{Br}(3, A)\cdots\text{Br}(3, C[z+1])$, 3.68 \AA ; $\text{Br}(3, A)\cdots\text{Br}(4, D[z+1])$, 3.81 \AA ; $\text{Br}(3, A)\cdots\text{Br}(4, A[z+1])$, 3.85 \AA .

Table 3 (cont.)

Crystal II

H K L				FO FC AC BC				H K L				FO FC AC BC				H K L				FO FC AC BC				H K L				FO FC AC BC													
8	2	4	641	537	931	88	6	7	4	217	150	190	7	5	0	5	935	964	954	133	6	5	5	1142	1124	1119	112	8	0	6	184	147	147	16	3	7	5	320	320	396	41

The P-Br bonds (mean value 2.162 Å) are equal within experimental error, as well as the angles Br-P-Br (mean 102.1°). Within the (N-P)₃ ring no significant differences are found between the P-N bonds (mean 1.576 Å), or between the angles N-P-N (mean 118.5°). However, there is a small, but significant, difference (3°) between the two independent angles P-N-P (119.3 and 122.4°). The smaller value of the angle P(2)-N(1)-P(2'), as compared with the angle P(2)-N(2)-P(1), together with the differences in the P-N bonds and the angles N-P-N (though the latter differences are not significant) describe a slight deformation of the six-membered ring of a type consistent with the bonding interaction, discussed above, between the nitrogen atom N(1) and the group P(1)-Br(1) of a second molecule

The (N-P)₃ ring is not entirely planar. The dihedral angles assignable to the P-N bonds are: -6.7° for N(1)-P(2), 6.3° for P(2)-N(2) and -5.5° for N(2)-P(1), indicating a slight chair-shaped conformation. The best plane through the six ring atoms, the phosphorus atoms having a weight four times that of the nitrogen atoms, is given by the equation

$$-0.3958 X + 0.9183 Z = 2.836 \quad (X \text{ and } Z \text{ in } \text{Å}).$$

The distances from the ring atoms to this plane and their standard deviations are ΔN(1) = -0.047 (0.011), ΔP(2) = 0.011 (0.003), ΔN(2) = -0.035 (0.008) and ΔP(1) = 0.008 (0.004) Å. The χ² test (Cruickshank & Robertson, 1953) on the planarity of the ring yielded χ² = 9990 and P < 0.001 for n = 6.

In the crystal structures of the related compounds N₃P₃Cl₆, N₃P₃Cl₅F (Olthof, 1969) and N₃P₃F₆ the

(N-P)₃ ring has been found to be planar within experimental error. The departure from planarity in the case of N₃P₃Br₆ is most likely due to intermolecular steric interactions, though the non-planarity is not apparent from the short intermolecular contacts, presented in Fig. 1 and 2.

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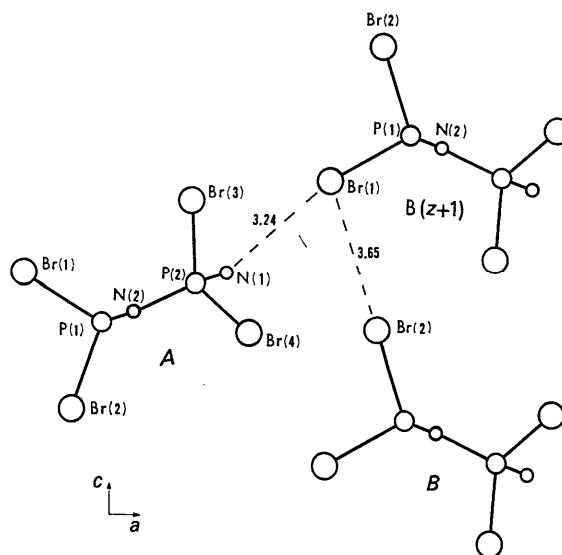
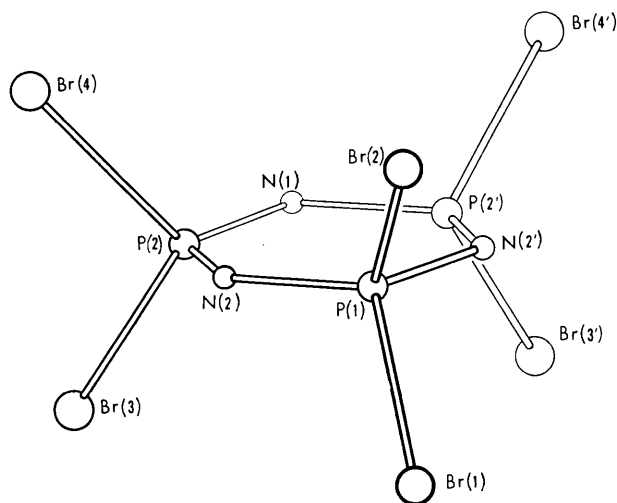


Fig. 2. Projection along the b axis of the molecules lying on the mirror plane *y* = 0.25 and short intermolecular distances in this plane.

Table 4. Bond lengths and valence angles with standard deviations

	Present work		Giglio & Puliti (1967)	
	Å	Å	Å	Å
N(1)-P(2)	1.588	0.006	1.55	0.035
N(2)-P(1)	1.565	0.008	1.60	0.033
N(2)-P(2)	1.575	0.008	1.60	0.045
P(1)-Br(1)	2.165	0.004	2.20	0.011
P(1)-Br(2)	2.158	0.004	2.17	0.017
P(2)-Br(3)	2.168	0.003	2.20	0.019
P(2)-Br(4)	2.159	0.003	2.14	0.013
P(2)-N(1)-P(2')	119.3°	0.6°	126.8°	2.6°
P(2)-N(2)-P(1)	122.4	0.5	118.6	2.4
N(2)-P(1)-N(2')	117.4	0.6	118.3	1.8
N(1)-P(2)-N(2)	119.0	0.4	115.8	2.0
Br(1)-P(1)-Br(2)	102.1	0.1	101.7	0.4
Br(3)-P(2)-Br(4)	102.0	0.1	103.4	0.4
N(1)-P(2)-Br(3)	106.9	0.5	106.7	2.4
N(1)-P(2)-Br(4)	108.7	0.4	111.9	1.7
N(2)-P(2)-Br(3)	109.3	0.4	114.4	1.8
N(2)-P(2)-Br(4)	109.7	0.4	104.0	1.8
N(2)-P(1)-Br(1)	109.1	0.4	113.4	1.7
N(2)-P(1)-Br(2)	109.0	0.4	103.8	2.1

Fig. 3. Molecule $N_3P_3Br_6$.

References

- BUSING, W. R. & LEVY, H. A. (1957). *Acta Cryst.* **10**, 180.
- COXON, G. E., SOWERBY, D. B. & TRANTER, G. C. (1965). *J. Chem. Soc.* p. 5697.
- CRUICKSHANK, D. W. J. & ROBERTSON, A. P. (1953). *Acta Cryst.* **6**, 698.
- CRUICKSHANK, D. W. J. (1961). *Computing Methods and the Phase Problem in X-ray Crystal Analysis*. Oxford: Pergamon Press.
- DOUGILL, M. W. (1963). *J. Chem. Soc.* p. 3211.
- DOYLE, P. A. & TURNER, P. S. (1968). *Acta Cryst.* **A24**, 390.
- GIGLIO, E. & PULITI, R. (1967). *Acta Cryst.* **22**, 304.
- HAZEKAMP, R., MIGCHELSEN, T. & VOS, A. (1962). *Acta Cryst.* **15**, 539.
- International Tables for X-ray Crystallography* (1962). Vol. III. Birmingham: Kynoch Press.
- MCGEACHIN, H. MCD. & TROMANS, F. R. (1961). *J. Chem. Soc.* p. 4777.
- OLTHOF, R. (1969). *Acta Cryst.* **B25**, 2040.
- PAULING, L. (1960). *The Nature of the Chemical Bond*. Ithaca: Cornell Univ. Press.
- SCHLUETER, A. W. & JACOBSON, R. A. (1968). *J. Chem. Soc.* **A**, p. 2317.
- WAGNER, A. J. & VOS, A. (1968). *Acta Cryst.* **B24**, 707.
- WIEBENGA, E. H. & KRACHT, D. (1969). *Inorg. Chem.* **8**, 738.
- WILSON, A. & CARROLL, D. F. (1960). *J. Chem. Soc.* p. 2548.
- ZACHARIASEN, W. H. (1967). *Acta Cryst.* **23**, 558.
- ZACHARIASEN, W. H. (1968). *Acta Cryst.* **A24**, 212.